



Chemistry

Name: _____

Section _____

MOLES AND MASS

Date: _____

Determine the number of moles in each of the quantities below.

1. 10. g of H_2O

2. 53.5 g of HNO_3

3. 75.6 g of $\text{C}_6\text{H}_{12}\text{O}_6$

4. 115 g of Hg_2Cl_2

5. 100. g of $\text{CaCl}_2 \cdot 6 \text{H}_2\text{O}$

Determine the number of grams in each of the quantities below.

1. 0.0035 moles of PbO_2

2. 0.641 moles of $\text{H}_2\text{C}_2\text{O}_4$

3. 10. moles of H_2SO_4

4. 1.79 moles of $\text{FeCl}_2 \cdot 4 \text{H}_2\text{O}$

5. 15 moles of $\text{C}_{12}\text{H}_{22}\text{O}_{11}$
