



# Chemistry

Name: \_\_\_\_\_

Section \_\_\_\_\_ Separating Mixtures WS Date: \_\_\_\_\_

List the physical properties involved and then explain how you might use them to your advantage in separating the following mixtures.

1. Sand and water \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_

2. Sugar and water \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_

3. Oil and water \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_

4. Sand and gravel \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_

5. A mixture of heptane (boiling point 98°C) and heptanol (boiling point 176°C) \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_

6. A mixture of iodine and sodium chloride (Hint: Iodine is not soluble in water.) \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_

7. A mixture of lead and aluminum pellets \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_

8. A mixture of salt and iron filings \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_